



Name..... Class.....

Make Sure You Have...

Iron (III) chloride (FeCl_3)

Potassium thiocyanate (KSCN)

Water (H_2O)

3 small beakers

1 measuring cylinder

2 microspatulas

What to Do...

1. Scoop a small mound of iron(III) chloride with a **clean** microspatula and place in a **clean** beaker.
2. Scoop a small mound of potassium thiocyanate with a **clean** microspatula and place in a **clean** beaker.

What colours are the Solids?

Iron(III) chloride.....

Potassium thiocyanate.....

What other form of iron does the iron(III) chloride remind you of?.....

3. Measure 20 ml of water into the measuring cylinder and pour it into the beaker containing the iron(III) chloride.
4. Repeat (3) for the beaker containing the potassium thiocyanate and stir each solution with a **clean** spatula to dissolve the solids.

What colours are the solutions?



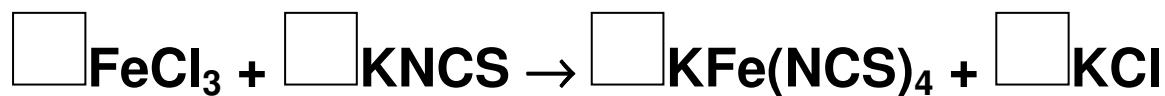
5. Pour the two beakers with solution into the empty beaker

What colour is the resulting solution?

Where in your body do you find iron?

What does your body use iron for?

The equation for this reaction is shown below. Please balance it.



6. Wash everything up (solutions can go down the sink).