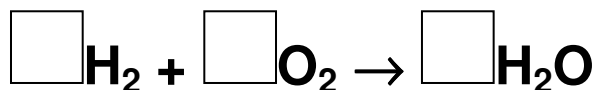




Name..... Class.....

1. The reaction you saw was hydrogen (H<sub>2</sub>) reacting with oxygen (O<sub>2</sub>) to make water (H<sub>2</sub>O).

The reaction equation is shown below. Can you balance it?



2. The bond enthalpies of H-H, O=O and H-O are 236 kJ mol<sup>-1</sup>, 497 kJ mol<sup>-1</sup> and 458 kJ mol<sup>-1</sup> respectively. How much energy is released in this reaction?

.....kJ mol<sup>-1</sup>

3. What 2 types of energy is the chemical energy converted into?

.....energy and .....energy

4. What is the name for something which speeds up a chemical reaction but is not changed by the reaction?

.....

5. Why does this reaction not happen immediately?.....

.....  
.....  
.....