

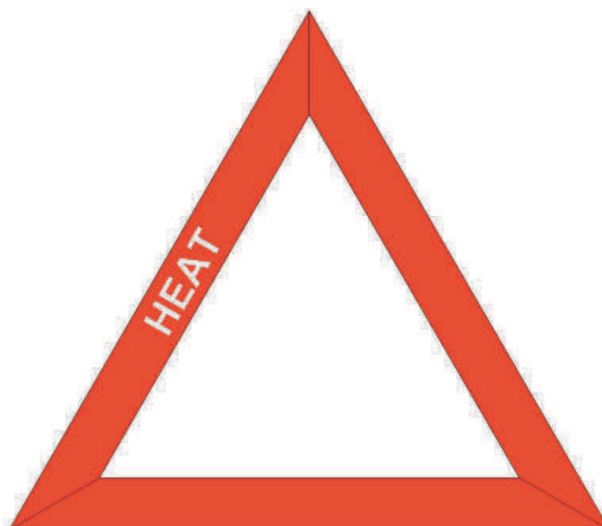


Name..... Class.....

- The reaction you saw was potassium permanganate ( $\text{KMnO}_4$ ) reacting with hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) to make oxygen ( $\text{O}_2$ ). The reaction equation is shown below. Can you balance it?



- The "Fire Triangle" is shown below. Please complete it with the three things you need for a fire to occur.



- You get a better fire here by adding potassium permanganate. Which of the three items on the fire triangle are you increasing?
- You can extinguish a fire by removing any one of the three sides of the fire triangle. If you use a carbon dioxide ( $\text{CO}_2$ ) fire extinguisher, which side of the fire triangle are you removing?
- When the potassium permanganate is added you see a violet flame due to the potassium ions. Name one other metal and the flame colour it gives.

Metal.....Colour.....